

***In-situ* characterization of the High-Pressure High-Temperature synthesis of tetragonal Na₂Si₃**

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Abstract

In a world of climate change and limited natural resources, the urgent need for sustainable energy technologies drives the search for novel functional materials made of abundant, non-toxic, and low-cost elements. In this context, sodium silicides are emerging as promising candidates for photovoltaics, energy storage, and thermoelectric applications. While the Na-Si phase diagram has been relatively unexplored, *in-situ* synchrotron techniques and high-pressure high-temperature (HP-HT) synthesis have enabled the search for additional phases. [1, 2]

At ambient pressure, NaSi is the only thermodynamically stable phase, but three metastable phases have also been reported: Type I clathrate Na₈Si₄₆ and type II clathrate Na_{x<24}Si₁₃₆ consists of a framework of Si cages that are filled with sodium ions. The third known structure, NaSi₆, was first discovered by Kurakevych *et al.* through *in-situ* monitoring of high-pressure high-temperature (HP-HT) synthesis. [1, 2, 3]

We report the discovery of a new Na₂Si₃ phase, formed under HP-HT conditions in the large volume press (LVP) located at the ID06-LVP beamline (ESRF) and identified using *in-situ* synchrotron X-ray diffraction. Na₂Si₃ was formed using a 10/5 multianvil assembly near 5 GPa at around 510 °C and is recoverable to ambient pressure and temperature. Additionally, we have explored different HP-HT synthesis paths for Na₂Si₃.

Na₂Si₃ adopts a tetragonal structure composed of two-dimensional layers of silicon, separated by sodium ions. The bonding can be rationalized within the Zintl-Klemm framework with a formal charge distribution (Na⁺)₂(Si⁰)(Si⁻)₂. *Ab initio* calculations confirmed this structure and predicted an indirect band gap of ~1.14 eV at ambient pressure, making Na₂Si₃ interesting for applications in photovoltaics and thermoelectricity. [4]

These findings highlight the importance of combining *in-situ* HP-HT experimentation with theoretical modelling to discover novel metastable materials.

[1] H. Morito, T. Yamada, T. Ikeda, H. Yamane, Journal of Alloys and Compounds. 2009, 480, 723-726.

[2] O. A. Kurakevych, T. A. Strobel, D. Y. Kim, T. Muramatsu, V. V. Struzhkin, Crystal Growth Des. 2013, 13, 303-307.

[3] C. Cros, M. Pouchard, P. Hagenmuller, J. Solid State Chem. 1970, 2-570-581.

[4] M. Redington, L. Schütte, D. O. A. Ali, D. Druzhbin *et al.*, submitted to Inorganic Chemistry, 2026.